**Science 10 Chemistry Day 9 – Ion Tiles with Polyatomics**

Write the formula of the compounds formed between the ions below:

a) Ni2+ and PO43– b) H+ and CO32–

c) Cr2+ and NO31– d) H+ and SO32–

e) Al3+ and SO42–

**Journal 3 (due tomorrow)**

1. What do we do differently when **naming** ionic compounds with **transition metals** (with more than one charge)? Give an example.
2. Would the correct name for **ScN** be **scandium (III) nitride**? Why or why not?
3. Explain how to determine **which ion** (which charge) of a transition metal is used when given the formula only. (example: PbS2)
4. Why are brackets needed if you need to indicate the presence of more than one polyatomic ion in a compound? Example: Why do we write Ca(OH)2 and not CaOH2?

**Evaluation**

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| --- | --- | --- | --- |
| **3** | **2** | **1** | **0** |
| Journal is thoughtful and demonstrates application and understanding of concepts investigated. | Journal is complete; however, application of concepts investigated is not always clear. | Journal is missing some components and/or does not demonstrate understanding of concepts investigated. | Not completed. |